



# FPCH001

## Chemistry 1

MQC3 Day 2014

*Macquarie City Campus*

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### Disclaimer

Macquarie University has taken all reasonable measures to ensure the information in this publication is accurate and up-to-date. However, the information may change or become out-dated as a result of change in University policies, procedures or rules. The University reserves the right to make changes to any information in this publication without notice. Users of this publication are advised to check the website version of this publication [or the relevant faculty or department] before acting on any information in this publication.

## General Information

Unit convenor and teaching staff

Lecturer in Charge and Tutor

David Johnson

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City Campus / St Andrew's Cathedral School

Contact lecturer

Credit points

3

Prerequisites

Corequisites

Co-badged status

Unit description

This unit is an introduction to chemistry at a level suitable for University entry. The study of chemistry involves learning about the nature of atoms, molecules and their interactions. Concepts covered include: classification and properties of elements and compounds, chemical reactions and equations, electro chemistry and an introduction to carbon chemistry. This unit will provide chemical learning using Australian educational methods. You will complete a significant amount of experimental work in the laboratory and will practice your English by participating in language based learning activities. The study of chemistry involves the students working individually and with others in practical and interactive media experiences related to the theoretical concepts considered in the course. It is expected that students studying the unit will begin to: apply investigative and problem-solving skills, effectively communicate the theoretical concepts considered in the course and appreciate the contribution that a study of chemistry makes to our understanding of the world.

## Important Academic Dates

Information about important academic dates including deadlines for withdrawing from units are available at <https://www.mq.edu.au/study/calendar-of-dates>

## Learning Outcomes

On successful completion of this unit, you will be able to:

Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding

Plan and perform laboratory investigations

Describe chemical changes in terms of substances involved and energy inputs and outputs

Describe and assess factors that influence the type and rate of chemical reactions

Relate the uses of carbon to the unique nature of carbon chemistry

Describe and predicts reactions involving carbon compounds

## General Assessment Information

### Missed Assessments

The only exception to not sitting an in-class test or examination at the designated time or handing in an assessment on the due date is because of a serious or unavoidable disruption.

Students who miss a formal assessment held in class or a final examination due to a serious and unavoidable disruption which commenced after the start of the study period must lodge a [Disruption to Studies](#) Notification via [ask.mq.edu.au](http://ask.mq.edu.au) within five (5) working days of the commencement of the disruption in order to apply for Special Consideration. The notification must be supported by appropriate [evidence](#).

In submitting a Disruption to Studies Notification, a student is acknowledging that they may be required to undertake additional work. The time and date, deadline or format of any required extra assessable work as a result of a Disruption to Studies Notification is not negotiable. Further, in submitting a Disruption to Studies Notification, a student is agreeing to make themselves available so that they can complete any extra work as required.

Students will be advised of the outcome of their [Disruption to Studies](#) Application via [ask.mq.edu.au](http://ask.mq.edu.au).

Please refer to the [Disruption to Studies Policy](#) for further details.

### Extensions & Late Submissions

To apply for an extension of time for submission of an assessment item, students must submit a notification of Disruptions to Studies via [ask.mq.edu.au](http://ask.mq.edu.au).

Grounds for extensions are usually serious illness, accident, disability, bereavement or other compassionate circumstances and must be substantiated with relevant evidence (e.g. professional authority form).

Late submissions without an approved extension will be penalised at a rate of **10% per day (weekend inclusive)**. This applies to assessments completed outside of class such as essays and assignments.

### Final Examinations and Final Assessment Tasks

Final exams and final assessments typically take place in Week 13 and the first 3 days of week 14. **Please note that you must pass the final exam or final assessment task in order to pass this unit.** You are expected to present yourself for examination at the time and place

designated in the Final Examination Timetable. Please note that no special consideration will be given to students who have booked flights out of the country prior to the conclusion of the examination period.

The Final Examination Timetable will be available in provisional form on the MQC Student Portal Noticeboard at <https://student.mqc.edu.au/NoticeBoard.htm> in approximately week 10 of this Session. You will have 1 week to give feedback to the Student Administration Manager should you have concerns or note any clashes in your final exam timetable. From week 12, you will also be able to view your personal final exam timetable via the [MQC Student Portal](#).

The examination timetable is produced to provide the maximum number of students with the least number of consecutive examinations. It is not uncommon for students of Macquarie University at both the City and North Ryde Campuses to be required to sit two consecutive examinations. A maximum of three consecutive exams is also permitted (for example, two on one day, and one the following morning). However, no student is required to sit four consecutive exams and if any student discovers their examination timetable contains four consecutive exams, they should immediately contact the [Student Administration Manager](#) to have an exam rescheduled.

Prior to the examination period, you should ensure that you are familiar with the [Examination Rules](#). You can find these under Exam Information on the MQC [Student Portal Noticeboard](#). A breach in any of these rules will lead to disciplinary action being undertaken.

**Students who miss a final exam or final assessment will be awarded a mark of 0 for the task and cannot pass the unit**, except for cases where a Disruption to Studies Notification is lodged and a Special Consideration is awarded. Please note that in submitting a Disruption to Studies Notification, a student is acknowledging that they may be required to undertake additional work. The time and date, deadline or format of any required extra assessable work as a result of a Disruption to Studies Notification is not negotiable.

## Supplementary Examinations

Supplementary final examinations are held during the scheduled Supplementary Final exam Period in the lead up to the subsequent teaching period.

Please note that results for supplementary exams may not be available until the conclusion of Week 2 of the subsequent teaching session and until supplementary results are released, continuing students may be prevented from enrolling in certain units in the subsequent teaching session.

Students in their final semester of study who undertake supplementary final exams should note that Formal Completion of the Foundation Program will not be possible until supplementary results are released and this may impact on their ability to enrol subsequent programs of study on time.

## Retention of Originals

It is the responsibility of the student to retain a copy of any work submitted and produce another copy of all work submitted if requested. Copies should be retained until after the release of final

results each Session.

In the event that a student is asked to produce another copy of work submitted and is unable to do so, they may be awarded zero (0) for that particular assessment task.

The University also reserves the right to request and retain the originals of any documentation/ evidence submitted to support notifications of disruptions to studies. Requests for original documentation will be sent to the applicant's University email address within six (6) months of notification by the student. Students must retain all original documentation for the duration of this six (6) month period and must supply original documents to the University within ten (10) working days of such a request being made.

## Turnitin

Students may be requested to submit assessments via Turnitin and in such instances any hard copies submitted without a Turnitin Report will not be marked.

Step by step guidance for Turnitin submissions can be found [here](#). Should you experience any difficulties with Turnitin submission, please see a Lab Demonstrator in Lab 311 at MQC.

If you experience difficulties submitting through Turnitin on the due date, you must email your work in electronic format to your lecturer using the email address provided in the unit guide. Late submissions will be penalised at 10% per day.

## Grading & Requirements to pass

This unit will use the following grading system:

- S – Satisfactory (50-100)
- F – Fail (0-49)

Grade descriptors and other information concerning grading are contained in the Macquarie University Grading Policy which is available at: <http://www.mq.edu.au/policy/docs/grading/policy.html>

To pass this unit, you must attempt all assessable components of the unit, pass the final exam and attain an overall mark of at least 50%. Failure to do so will result in an F (fail) grade being recorded.

## Provision of Feedback

Marks awarded for assessment items will also be available on the MQC Student Portal within fourteen (14) days of the due date. Typically you will also be able to collect your submitted assignments in this same timeframe.

If you wish to receive further feedback from your instructor, you should contact them directly using the contact details provided in this guide.

Students may seek general feedback about their performance in a unit up to 6 months following results release.

## Contacting Staff and Getting Help

Foundation students may approach teaching staff for one-on-one help in one of three ways:

- During Consultation sessions. For details about consultation sessions and Consultation times, please refer to timetable provided on the Macquarie [City Campus Portal Noticeboard](#).
- Using the "Questions for your instructor" dialogue provided in Week 0 of the respective unit in [iLearn](#).
- Using the instructor's email address provided in the Unit Guide of the respective unit.

For all university related correspondence, students are required to use their official MQ student email account which may be accessed via the [Macquarie University Student Portal](#). Enquiries from personal email accounts will not be replied to.

## Assessment Tasks

Name	Weighting	Due
<a href="#">Weekly Quizzes</a>	15%	Weeks 1-12
<a href="#">Class Test 1</a>	15%	Week 4
<a href="#">Class Test 2</a>	35%	Week 10
<a href="#">Final Examination</a>	35%	Examination Period

## Weekly Quizzes

Due: **Weeks 1-12**

Weighting: **15%**

A brief quiz will be held at the start of most classes. It will be a mixture of written and oral questions based on material covered in class in the previous week(s). The first quiz is worth 0%. It is mainly to give the lecturer an idea of your current knowledge. All other quizzes will be worth 1% each. The best fifteen quiz marks will be included in the final mark (15%).

On successful completion you will be able to:

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions

- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

## Class Test 1

Due: **Week 4**

Weighting: **15%**

This practical test on which students will work individually or in a small group to carry out an experimental task (5%) as well as an individual written test paper covering work from weeks 1-3 (10%).

On successful completion you will be able to:

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe and predicts reactions involving carbon compounds

## Class Test 2

Due: **Week 10**

Weighting: **35%**

This will be a written examination paper covering work from weeks 1-9 (35%).

On successful completion you will be able to:

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions

## Final Examination

Due: **Examination Period**

Weighting: **35%**

The final examination is of 2 hours duration and will consist of a written examination covering work from weeks 1-13 (25%) and a practical test in which students will work individually or in a small group to carry out an experimental task (10%). The final exam will be held during the final examination period in either Week 13 or 14. **Please note that you must pass the final exam in order to pass this unit.**

On successful completion you will be able to:

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

## Delivery and Resources

### Classes

Weekly contact will be 5 hours consisting 2 x 2.5 hour lessons.

Classes will take place at St Andrew's Cathedral School, which is located near Town Hall station (483 George St, Sydney NSW 2000).

The lecturer will meet students at MQC in Week 1 (in the room indicated on the students' timetable) and take them down to St Andrew's. Students will be expected to make their way to St Andrew's from thereon. Please make sure you allow sufficient time to make your way to the School. If you have a class immediately prior to your Science lesson at St Andrew's, you are permitted to leave that class 15 minutes early in order to arrive at Sta Andrew's on time.

While at St Andrew's, students will be expected to comply with the school's rules and procedures.

During lessons, new content will be presented and explained by the lecturer. During laboratory work and tutorials participants will have more opportunities to engage in discussion and activities. (This unit will not comprise any consultation sessions.) Attendance of all sessions is compulsory.

Timetables can be found on the Noticeboard on the [City Campus Student Portal](#).

If any scheduled class falls on a public holiday a make-up lesson may be scheduled, usually on a Saturday. Scheduled make-up days are noted in the Teaching Schedule of the Unit Guide and attendance is compulsory. Where appropriate the instructor may require students to complete alternative activities on-line rather than attending a make-up lesson.

### Learning and Teaching Activities

This unit will be taught will require students to complete pre-set practical exercises based on material discussed in lectures and tutorials. Students will be required to work independently as well as in small groups and engage in class discussions.

It is expected that all students purchase the prescribed text and read in advance to ensure that they are well prepared for the content covered in each lecture. It is recommended that:



- Before commencing the prescribed chapter readings in the Unit Schedule each week, Students study “Important new terms” listed at the end of each chapter. This is a summary of the important new terms found in each chapter. Page references are provided for each new term relevant to the chapter readings.
- Students complete the “exercises” for each chapter reading in the Unit Schedule each week.
- Students complete the “Extended Response Exam-style Questions” at the end of each module, for questions relevant to the Unit Schedule chapter readings (see Contents page v)

iLearn will also be used to post lecture and tutorial materials and also communicate with students so it is expected that students will check this resource on a regular basis.

## iLearn

[iLearn](#) is Macquarie's online learning management systems. The following unit specific information will be available on the website:

- Announcements
- Staff contact details
- Lecture notes and recordings
- Learning and teaching activities and resources
- Assessment information
- Tutorial questions and solutions
- Assessment submission tools such as Turnitin
- Other relevant material

Please note that you must enrol in a unit via [eStudent](#) in gain access to the unit in iLearn.

You are required to regularly check the website and use it as an information and resource centre to assist with your learning.

Ensure that when you have finished using the website, you log out. Failure to do so could allow unauthorised access to your account.

Please contact the IT helpdesk (Ph. 02 9850 4357) or lodge a ticket using [OneHelp](#) if you need assistance accessing iLearn.

## Required and Recommended Texts and Materials

- **Prescribed textbook(s):**

Weeks 1-9: *Conquering Chemistry Preliminary Course*, Roland Smith, 4<sup>th</sup> edition.  
Publisher McGraw Hill. ISBN: 9780170226790

Weeks 10-13: *Conquering Chemistry HSC Course*, Roland Smith, 4<sup>th</sup> edition. Publisher McGraw Hill. ISBN 9780170226820. (also used in Chemistry 2)

N.B. Older editions of these texts are also suitable.

- **Recommended textbook(s):**

*Complete Chemistry*, Ingham and Gallagher. Oxford university press. ISBN 9780199138784

All prescribed textbooks will be made available to students to purchase at the Phillip Street Coop Bookshop.

Students can view a full list of textbooks for all units on the Macquarie [City Campus Student Portal Noticeboard](#).

- **Technology Used and Required**

- Access to computer and internet outside class
- Access to [iLearn](#)

## Unit Schedule

Week	Topic	Readings
Beginning:		
<b>Week 1</b>	<b>Topic 1 – Chemical Earth</b>	<b>Preliminary Text</b>
Mon 20 October	<ul style="list-style-type: none"><li>• Mixtures and pure substances</li><li>• Elements and compounds</li><li>• Separation of mixtures</li><li>• Properties for identifying substances</li><li>• Gravimetric analysis</li><li>• Metals and non-metals</li><li>• Physical properties and uses of metals</li><li>• The Periodic Table as a list of elements</li><li>• Particle nature of matter</li><li>• Particle theory of solids, liquids and gases</li><li>• Atoms and molecules</li><li>• Symbols and formulae</li></ul> <p><b>Assessment 1 begins - quizzes continue every class</b></p>	Chapter 1(1.1-1.2, 1.5-1.20, 1.23-1.25)  Chapter 2 (2.1-2.6)

<b>Week 2</b>  Mon 27 October	<ul style="list-style-type: none"> <li>· Nucleus and electron cloud</li> <li>· Structure of some simple atoms</li> <li>· Atomic number and mass number</li> <li>· Energy levels for electrons</li> <li>· Electron configurations</li> <li>· Stable electron configurations</li> <li>· The Periodic Table</li> <li>· Achieving noble gas configurations</li> <li>· Formation of ions, ionic bonding and the Periodic Table</li> <li>· Covalent bonding</li> <li>· Covalency and the Periodic Table</li> <li>· Electron-dot structures for ions</li> <li>· Properties of ionic and covalent</li> <li>· Properties of solids summarized</li> <li>· Physical and chemical changes</li> <li>· Decomposition and direct combination reactions</li> </ul>	Chapter 2 (2.7-2.22)  Chapter 3 (3.1-3.3)
<b>Week 3</b>  Mon 3 November	<b>Topic 2 – Metals</b> <ul style="list-style-type: none"> <li>· Energy changes for these reactions</li> <li>· Chemical equations</li> <li>· Formulae and names for ionic compounds</li> <li>· Formulae and names for binary covalent compounds</li> <li>· Reactions of metals with oxygen, water and dilute acids</li> <li>· An activity series for metals and its use in selecting metals for particular purposes</li> </ul>	Chapter 3 (3.4-3.12)  Chapter 4 (4.9-4.11 )
<b>Week 4</b>  Mon 10 November	<ul style="list-style-type: none"> <li>· Ionization energy</li> <li>· Why quantitative calculations are necessary Relative atomic mass (atomic weight)</li> <li>· Relative molecular mass (molecular weight)</li> <li>· The mole &amp; converting between masses</li> <li>· Moles and numbers of atoms and molecules</li> <li>· Per cent composition</li> </ul> <b>Assessment 2 – Practical test and Written Quiz</b>	Chapter 4 (4.12-4.13)  Chapter 5 (5.1 – 5.6)
<b>Week 5</b>  Mon 17 November	<ul style="list-style-type: none"> <li>· Calculating formulae for compounds</li> <li>· Empirical and molecular formulae</li> <li>· Chemical equations and moles</li> <li>· Mass calculations using equation</li> <li>· Yield and its importance for extracting metals</li> <li>· The contributions of Dalton</li> <li>·</li> </ul>	Chapter 5 (5.7 – 5.13)

<b>Week 6</b>  Mon 24 November	<ul style="list-style-type: none"> <li>· Brief historical development of the Periodic table</li> <li>· Ionisation</li> <li>· Energy</li> <li>· Valence and position in the Periodic Table</li> <li>· Electronegativity</li> </ul>	Chapter 6 (6.1,6.3,6.4, 6.5), electronegativity graph (page 166 )  ionisationenergy graph page161
<b>Week 7</b>  Mon 1 December	<b>Topic 3 - Water</b> <ul style="list-style-type: none"> <li>· Solutions and solubility rules</li> <li>· Precipitation reactions</li> <li>· Concentration of solutions (Molarity only)</li> <li>· Molarity</li> <li>· Equilibrium in solutions</li> </ul>	Chapter 7(,7.1-7.3, 7.8 -7.10,7.12,  Chapter 8 (8.1-8.11)
<b>Week 8</b>  Mon 8 December	<b>Topic 4 – Energy</b> <ul style="list-style-type: none"> <li>· Fossil fuels</li> <li>· Double and triple covalent bonds</li> <li>· Carbon–carbon bonds</li> <li>· Fractional distillation of crude oil</li> <li>· Alkanes</li> <li>· Alkenes</li> <li>· Functional groups</li> </ul>	Chapter 9 (9.3, 9.6-9.11)
<b>Week 9</b>  Mon 15 December	<ul style="list-style-type: none"> <li>· Exothermic and endothermic reactions</li> <li>· Enthalpy—its meaning and use</li> <li>· Measuring enthalpy changes for reactions</li> <li>· Heat of combustion</li> <li>· Pollution from burning fossil fuels</li> <li>· Greenhouse effect</li> <li>· Rates of combustion reactions</li> </ul>	Chapter 8 (8.10, 8.11, 8.12)  Chapter 10 (10.1-10.4, 10.8-10.12)
<b>Week 10</b>  Mon 22 December	<b>Assessment 3 – Written Test</b>  <b>Topic 5 - Production of Materials</b> <ul style="list-style-type: none"> <li>· Hydrocarbons and their reactions</li> </ul>	<b>HSC Text</b>  Chapter 1 (1.1- 1.6), page 17
<b>Week 11</b>  Mon 5 January	<ul style="list-style-type: none"> <li>· Polymers</li> <li>· Ethanol and its reactions</li> <li>· Electrochemistry</li> </ul>	Chapter 1 (1.6-1.7) (1.16- 1.20), page 17  Chapter 2 (2.1-2.2, 2.4-2.8, 2.10-211)

<b>Week 12</b>	· Electrochemistry continued	Chapter 2 (2.13 -2.14)
Mon 12 January	· Nuclear Chemistry	Chapter 3 (3.1,3.2,3.4,3.5,3.9)
<b>Week 13</b>	· Revision	
Mon 19 January	<b>Assessment 4 – Final exam</b>  (Final Exams may be Held in Week 13 or 14, during the scheduled final exam period. Please refer to the Information Provided on the Portal Noticeboard). <b>Please note that you must pass the final exam in order to pass this unit.</b>	

## Other Important Dates

Public holidays & make-up days  New Year's Day Make-up: Saturday 20 December  (Please note that online lessons may be organised in lieu of make-up day).
Session Break  Thursday 25 December 2014 – Thursday 1 January 2015 inclusive
Exam Period:  Monday 19 January 2015 – Wednesday 28 January 2015 (Sat & Sun excluded)  Please note all final exams will be held in the City Campus.
Results Release:  Session 3 2014 results are scheduled to be released to students via e-Student and MQC Student Portal  on Friday 13 February 2015

## Policies and Procedures

Macquarie University policies and procedures are accessible from [Policy Central](#). Students should be aware of the following policies in particular with regard to Learning and Teaching:

Academic Honesty Policy [http://mq.edu.au/policy/docs/academic\\_honesty/policy.html](http://mq.edu.au/policy/docs/academic_honesty/policy.html)

Assessment Policy <http://mq.edu.au/policy/docs/assessment/policy.html>

Grading Policy <http://mq.edu.au/policy/docs/grading/policy.html>

Grade Appeal Policy <http://mq.edu.au/policy/docs/gradeappeal/policy.html>

Grievance Management Policy [http://mq.edu.au/policy/docs/grievance\\_management/policy.html](http://mq.edu.au/policy/docs/grievance_management/policy.html)

Disruption to Studies Policy [http://www.mq.edu.au/policy/docs/disruption\\_studies/policy.html](http://www.mq.edu.au/policy/docs/disruption_studies/policy.html) *The Disruption to Studies Policy is effective from March 3 2014 and replaces the Special Consideration Policy.*

In addition, a number of other policies can be found in the [Learning and Teaching Category](#) of Policy Central.

## Student Code of Conduct

Macquarie University students have a responsibility to be familiar with the Student Code of Conduct: [https://students.mq.edu.au/support/student\\_conduct/](https://students.mq.edu.au/support/student_conduct/)

## Academic Honesty

The nature of scholarly endeavour, dependent as it is on the work of others, binds all members of the University community to abide by the principles of academic honesty. Its fundamental principle is that all staff and students act with integrity in the creation, development, application and use of ideas and information. This means that:

- all academic work claimed as original is the work of the author making the claim
- all academic collaborations are acknowledged
- academic work is not falsified in any way
- when the ideas of others are used, these ideas are acknowledged appropriately.

Further information on the academic honesty and schedule of penalties that will apply to breaches please consult the [Academic Honesty Policy](#).

If you are unsure about how to incorporate scholarly sources into your own work, please speak to your Instructor or the Student Services team well in advance of your assessment. You may also enrol in [StudyWise](#) or visit the University's [Library Webpage](#) for more resources.

## Final Examination Script Viewings and Grade Appeals

If, at the conclusion of the unit, you have performed below expectations, and are considering lodging an appeal of grade and/or viewing your final exam script please refer to [http://www.city.mq.edu.au/new\\_and\\_current\\_students/appeals/](http://www.city.mq.edu.au/new_and_current_students/appeals/) for information about associated cut off dates.

Please note that any requests to view exam papers must be booked in immediately following results release.

Before submitting a Grade Appeal, please ensure that you read the [Grade Appeal Policy](#) and noted valid grounds for appeals.

## Attendance

Please refer to the [Attendance Policy for Foundation Students](#).

A minimum level of 80% attendance is compulsory for all classes, including consultation sessions and any make-up classes scheduled on weekends. Attendance will be recorded in every lesson and note made of any lateness or period of absence from class.

Where a student is present for only for a minor portion of a lesson (for example arrives late, leaves early, leaves the class frequently or for lengthy periods, engages in inappropriate or unrelated activities or does not participate actively in the majority of the lesson) the instructor reserves the right to mark a student absent for that particular lesson and make note of such incidents.

Students should note that absenteeism (including partial absenteeism) not only has a negative impact on not only their overall attendance record and their academic progress, but could also have ramifications for their visas or eligibility for social benefits where relevant.

In cases of unavoidable non-attendance due to illness or circumstances beyond control, students are advised to lodge a [Disruption to Studies](#) Notification via [ask.mq.edu.au](http://ask.mq.edu.au) even if they have not missed a formal assessment task so that appropriate records of the reasons for unavoidable attendance can be made on their record.

## Course Progression

Macquarie City Campus monitors Foundation students' course progress. Please refer to the [Course Progress Policy](#).

To maintain satisfactory program performance students are required to pass 50% or more of their enrolled units in each session.

Students who fail to make satisfactory course progress will be classified as "at risk" students and may have conditions placed upon their enrolment.

International students must comply with the Course Progress policy in order to meet the conditions of their visa.

## Student Support

Macquarie University provides a range of support services for students. For details, visit <http://students.mq.edu.au/support/>

## Learning Skills

Learning Skills ([mq.edu.au/learningskills](http://mq.edu.au/learningskills)) provides academic writing resources and study strategies to improve your marks and take control of your study.

- [Workshops](#)
- [StudyWise](#)
- [Academic Integrity Module for Students](#)
- [Ask a Learning Adviser](#)

## Student Support at Macquarie City Campus

Macquarie City Campus students who require assistance or support are encouraged to contact Student Services ([studentadvisor@city.mq.edu.au](mailto:studentadvisor@city.mq.edu.au)) or make an appointment to see a student advisor at Reception on Level 2.

Macquarie University Campus Wellbeing services are also available at the City Campus. If you

would like to make an appointment, please email [info@city.mq.edu.au](mailto:info@city.mq.edu.au) or visit their website at: <http://www.campuslife.mq.edu.au/campuswellbeing>.

## Academic Support at Macquarie City Campus

Macquarie city campus provides free tutoring / support classes to its student. Support is available for Accounting, numeracy and essay and report writing, research presentation and referencing skills.

Students who are experiencing difficulties in these areas are advised to attend these classes on a drop-in basis. So that the tutor can assist best, students must bring the work (e.g. assignment draft, essay draft, homework problem) with which that they are having difficulties.

For further information about tutoring services, please refer to the [City Campus Portal Noticeboard](#) under Timetables, Tutor Availability.

If you require additional support with university skills, you may also consider enrolling in [UNIWIS E](#). UNIWIS E is an iLearn resource which provides:

- Online learning resources and academic skills workshops
- What is expected of you as a student at Macquarie University
- Personal assistance with your learning & study related questions
- Key strategies and tips that you can use to achieve successful learning both in and out of the classroom
- The definitions and examples of the types of assignments you will encounter in your units

Additional study spaces are also available on Level 1.

## Student Services and Support

Students with a disability are encouraged to contact the [Disability Service](#) who can provide appropriate help with any issues that arise during their studies.

## Student Enquiries

For all student enquiries, visit Student Connect at [ask.mq.edu.au](http://ask.mq.edu.au)

## IT Help

For help with University computer systems and technology, visit <http://informatics.mq.edu.au/help/>.

When using the University's IT, you must adhere to the [Acceptable Use Policy](#). The policy applies to all who connect to the MQ network including students.

If you need IT support with any of the Macquarie University Systems please see <http://informatics.mq.edu.au/help/>, [lodge a One Help](#) ticket or call 02 9850-4357.



Students must use their Macquarie University email addresses to communicate with staff as it is University policy that the University issued email account is used for official University communication.

## IT Help at Macquarie City Campus

A lab demonstrator is situated in Lab 311 and can help you with any usage of university systems or resetting your password.

You may also refer to the Online Systems Password Document which has been made available on the [City Campus Student Portal Noticeboard](#).

Whilst utilising the City Campus IT facilities, students are expected to act responsibly. The following regulations apply to the use of computing facilities and online services:

- Accessing inappropriate web sites or downloading inappropriate material is not permitted.
- Material that is not related to coursework for approved unit is deemed inappropriate.
- Downloading copyright material without permission from the copyright owner is illegal, and strictly prohibited. Students detected undertaking such activities will face disciplinary action, which may result in criminal proceedings.

Non-compliance with these conditions may result in disciplinary action without further notice.

## Equipment available for loan

Students may borrow headphones for use in the Macquarie City Campus computer labs (210, 307, 311, 608) or a video recorder.

Please ask at Level 2 Reception for details. You will be required to provide your MQC Student ID card which will be held as a deposit while using the equipment.

## Graduate Capabilities

### Capable of Professional and Personal Judgement and Initiative

We want our graduates to have emotional intelligence and sound interpersonal skills and to demonstrate discernment and common sense in their professional and personal judgement. They will exercise initiative as needed. They will be capable of risk assessment, and be able to handle ambiguity and complexity, enabling them to be adaptable in diverse and changing environments.

This graduate capability is supported by:

### Learning outcome

- Plan and perform laboratory investigations

### Assessment tasks

- Weekly Quizzes

- Class Test 1
- Class Test 2
- Final Examination

## Discipline Specific Knowledge and Skills

Our graduates will take with them the intellectual development, depth and breadth of knowledge, scholarly understanding, and specific subject content in their chosen fields to make them competent and confident in their subject or profession. They will be able to demonstrate, where relevant, professional technical competence and meet professional standards. They will be able to articulate the structure of knowledge of their discipline, be able to adapt discipline-specific knowledge to novel situations, and be able to contribute from their discipline to inter-disciplinary solutions to problems.

This graduate capability is supported by:

### Learning outcomes

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

### Assessment tasks

- Weekly Quizzes
- Class Test 1
- Class Test 2
- Final Examination

## Critical, Analytical and Integrative Thinking

We want our graduates to be capable of reasoning, questioning and analysing, and to integrate and synthesise learning and knowledge from a range of sources and environments; to be able to critique constraints, assumptions and limitations; to be able to think independently and systemically in relation to scholarly activity, in the workplace, and in the world. We want them to have a level of scientific and information technology literacy.

This graduate capability is supported by:

### Learning outcomes

- Explain trends and relationships between elements in terms of atomic structure, the

periodic table and bonding

- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

## Assessment tasks

- Weekly Quizzes
- Class Test 1
- Class Test 2
- Final Examination

## Problem Solving and Research Capability

Our graduates should be capable of researching; of analysing, and interpreting and assessing data and information in various forms; of drawing connections across fields of knowledge; and they should be able to relate their knowledge to complex situations at work or in the world, in order to diagnose and solve problems. We want them to have the confidence to take the initiative in doing so, within an awareness of their own limitations.

This graduate capability is supported by:

## Learning outcomes

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Plan and perform laboratory investigations
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

## Assessment tasks

- Weekly Quizzes
- Class Test 1
- Class Test 2
- Final Examination

## Effective Communication

We want to develop in our students the ability to communicate and convey their views in forms effective with different audiences. We want our graduates to take with them the capability to read, listen, question, gather and evaluate information resources in a variety of formats, assess, write clearly, speak effectively, and to use visual communication and communication technologies as appropriate.

This graduate capability is supported by:

### Learning outcomes

- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry

### Assessment tasks

- Weekly Quizzes
- Class Test 2
- Final Examination

## Engaged and Ethical Local and Global citizens

As local citizens our graduates will be aware of indigenous perspectives and of the nation's historical context. They will be engaged with the challenges of contemporary society and with knowledge and ideas. We want our graduates to have respect for diversity, to be open-minded, sensitive to others and inclusive, and to be open to other cultures and perspectives: they should have a level of cultural literacy. Our graduates should be aware of disadvantage and social justice, and be willing to participate to help create a wiser and better society.

This graduate capability is supported by:

### Learning outcome

- Describe and predicts reactions involving carbon compounds

### Assessment tasks

- Weekly Quizzes
- Class Test 1
- Final Examination

## Socially and Environmentally Active and Responsible

We want our graduates to be aware of and have respect for self and others; to be able to work with others as a leader and a team player; to have a sense of connectedness with others and country; and to have a sense of mutual obligation. Our graduates should be informed and active

participants in moving society towards sustainability.

This graduate capability is supported by:

## Learning outcomes

- Explain trends and relationships between elements in terms of atomic structure, the periodic table and bonding
- Describe chemical changes in terms of substances involved and energy inputs and outputs
- Describe and assess factors that influence the type and rate of chemical reactions
- Relate the uses of carbon to the unique nature of carbon chemistry
- Describe and predicts reactions involving carbon compounds

## Assessment tasks

- Weekly Quizzes
- Class Test 1
- Class Test 2
- Final Examination

## Changes since First Published

Date	Description
28/10/2014	To pass this unit, you must attempt all assessable components of the unit, pass the final exam and attain an overall mark of at least 50%. Failure to do so will result in an F (fail) grade being recorded.